UNIDADES QUÍMICAS

de concentración

molaridad

$$M = \frac{\text{moles de soluto}}{\text{litros de solución}}$$

$$M = \frac{m}{(PM)(V)}$$

molalidad

$$m = \frac{\text{moles de soluto}}{\text{Kg de solvente}}$$

$$m = \frac{\text{gramos}}{(PM)(\text{Kg agua})}$$

normalidad

$$N = \frac{\text{equivalente gr soluto}}{\text{litro / de solución}}$$

Ejercicios molaridad

1. $M = \frac{m}{(PM)(V)}$ $\quad \frac{82,5g}{(46)(0,45l)} = 20,7$

$M = 3,98$

2. $M = \frac{478}{7} = 0,68$ $\quad V = 7000/1000 = 7$

3 $m = 3,65\ g$
$V = 2,00l$

$H = 1 \times 1 = 1$
$Cl = 1 \times 35 = 35$
$36\ g/mol$

$M = \frac{3,65}{(36)(2,00)} = \frac{3,65}{72} =$

$= 0,05$

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4 m = 49,09 g

V = 0,25 l

H = 2 x 1 = 2

S = 1 x 32 = 32

O = 4 x 16 = 64

96 g/mol

$$M = \frac{19,04}{(96)\cdot(0,25)} = \frac{49,04}{24} = 2,04$$

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