

$$1) \quad C = \frac{40}{12} = 3,3 = 3,3 / 3,3 = 1$$

$$H = \frac{6,79}{1} = 6,7 = 6,7 / 3,3 = 2$$

$$O = \frac{53,3}{16} = 3,3 = 3,3 / 3,3 = 1$$

$= C_1 H_2 O_1$

$$C = 12 \times 1 = 12$$

$$H = 1 \times 2 = 2$$

$$O = 16 \times 1 = 16$$

30

$$n = \frac{90}{30} = 3 = C_1 H_2 O_1 \rightarrow C_3 H_6 O_3$$

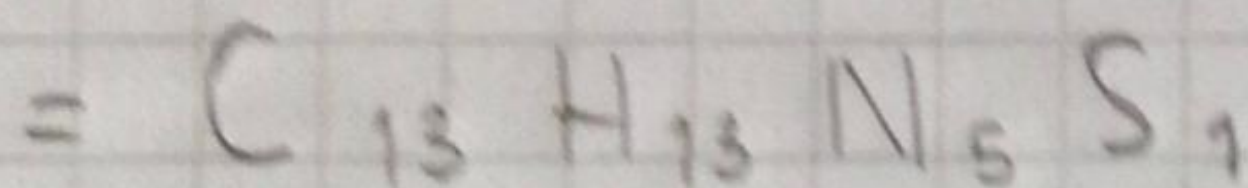
2) Obtenga la fórmula empírica y molecular si tiene 48% carbono, Hidrogeno 4%, Nitrogeno 22,4%, Azufre 12,8%, Oxigeno 12,8% y masa de 1000 g.

$$C = \frac{48 \text{ g}}{12 \text{ g/mol}} = 4 = 4 / 0,3 = 13$$

$$H = \frac{4 \text{ g}}{1 \text{ g/mol}} = 4 = 4 / 0,3 = 13$$

$$N = \frac{22,4 \text{ g}}{14 \text{ g/mol}} = 1,6 = 1,6 / 0,3 = 5$$

$$S = \frac{12,8 \text{ g}}{32 \text{ g/mol}} = 0,3 = 0,3 / 0,3 = 1$$



$$C = 12 \times 13 = 156$$

$$H = 1 \times 13 = 13$$

$$N = 14 \times 5 = 70$$

$$S = 32 \times 1 = 32$$

$$271$$

$$n = \frac{10000 \text{ g}}{271} = 4, \text{ } \Rightarrow \text{C}_{13} \text{H}_{13} \text{N}_5 \text{S}_1$$

