

REDOX

Propósito:
analizar y comprender el balanceo de
ecuaciones por redox o oxidoreducción

número de oxidación



Oxidación:

aumenta el número de oxidación



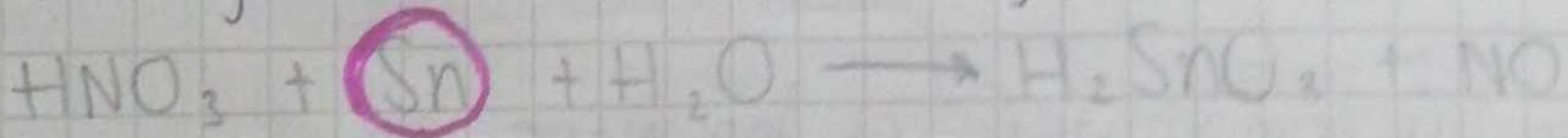
Reducción:

Disminuye el número de oxidación

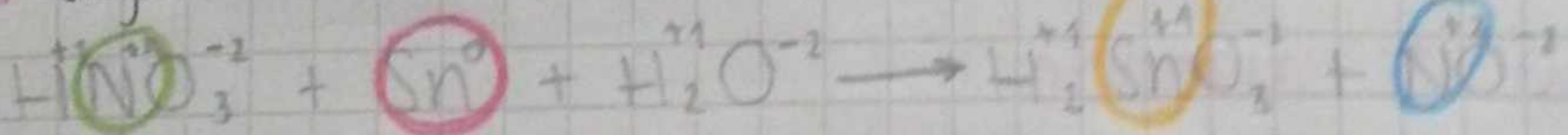


• Hidrogenos = +1

• Oxigeno = -2



1. asignar estados de oxidación



Si el elemento esta solo su estado de oxidación siempre sera 0.

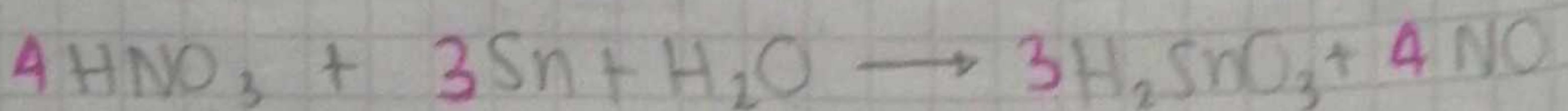
2. Identificar que elemento se oxido y cual se redujo y en cuanto.



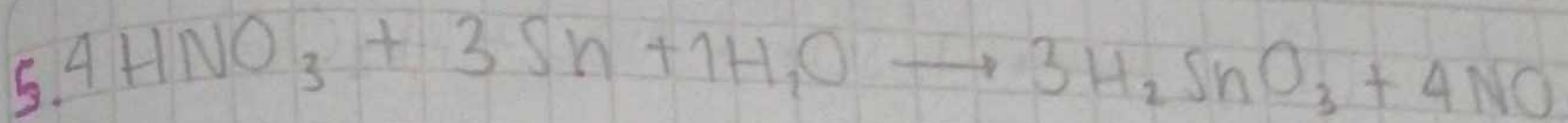
3. Cambie la diferencia entre estos



4. Ubique estos números en la ecuación



5. Verifique que este balanceada



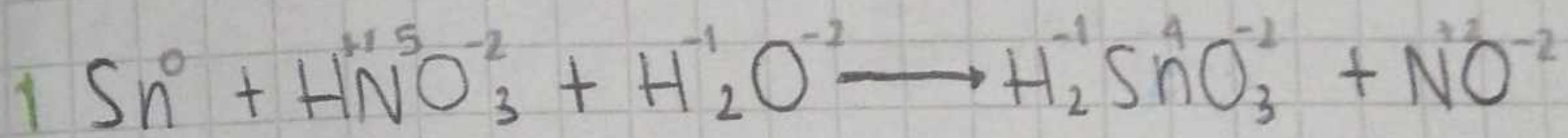
4 N 4

3 Sn 3

6 H 6

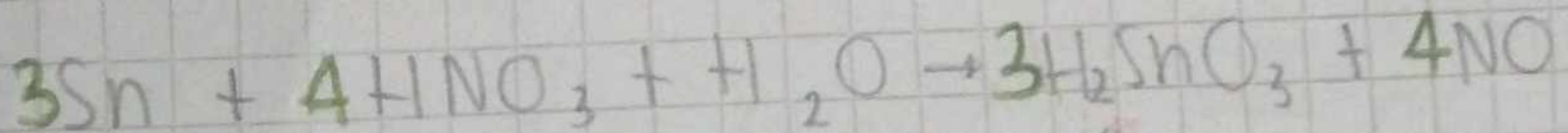
13 O 13

ACTIVIDAD



$\text{N}^{+5} \longrightarrow \text{N}^{+2}$ (3) reducción

$\text{Sn}^0 \longrightarrow \text{Sn}^{+4}$ (4) oxidación

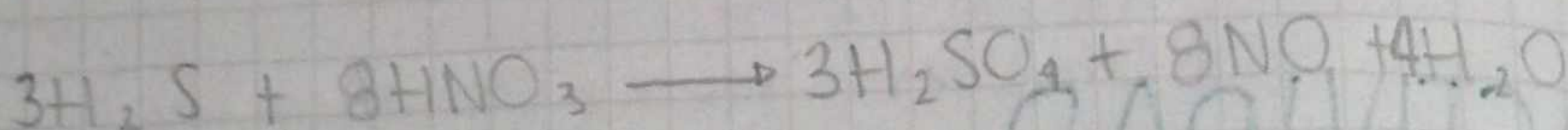
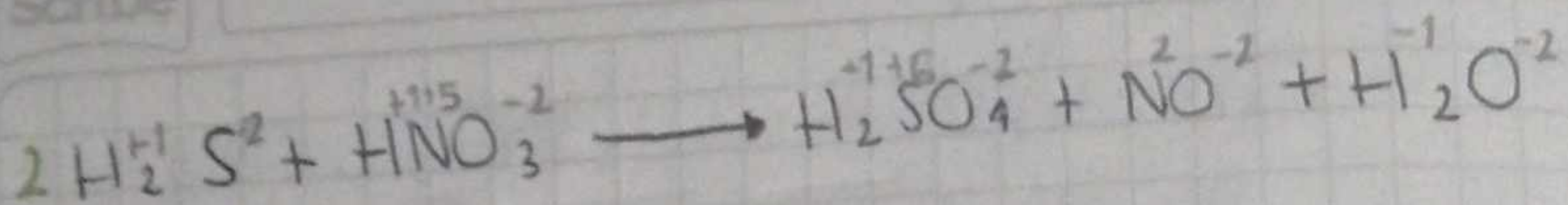


3 Sn 3

4 N 4

6 H 6

13 O 13

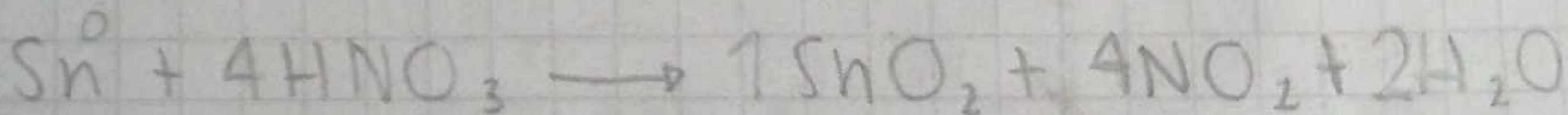
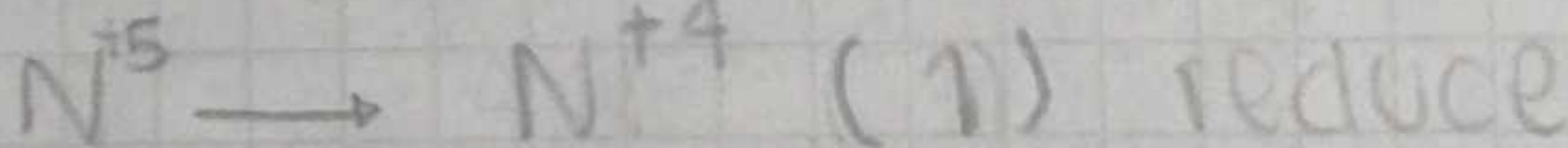
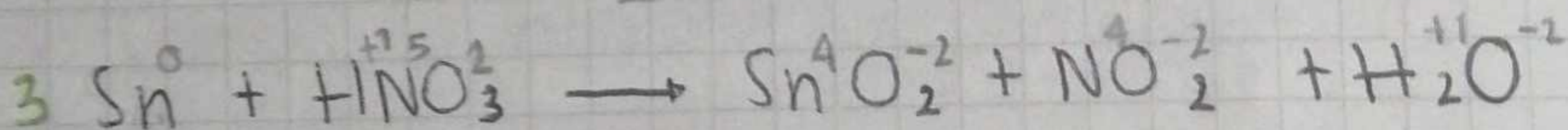


3 S 3

8 N 8 O₂ H + ~~8 H₂O~~

14 H 14

24 O 24

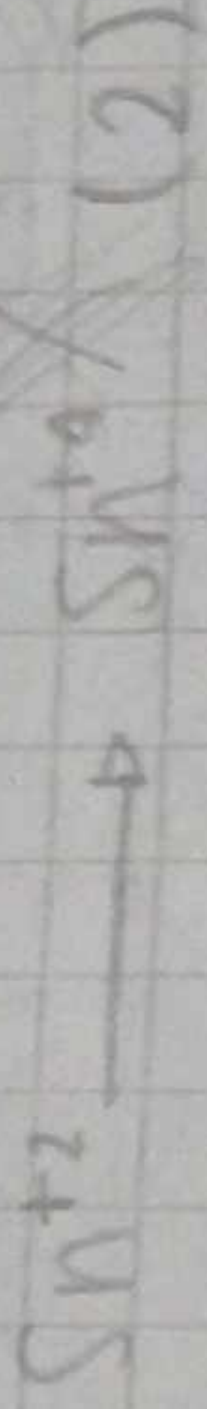
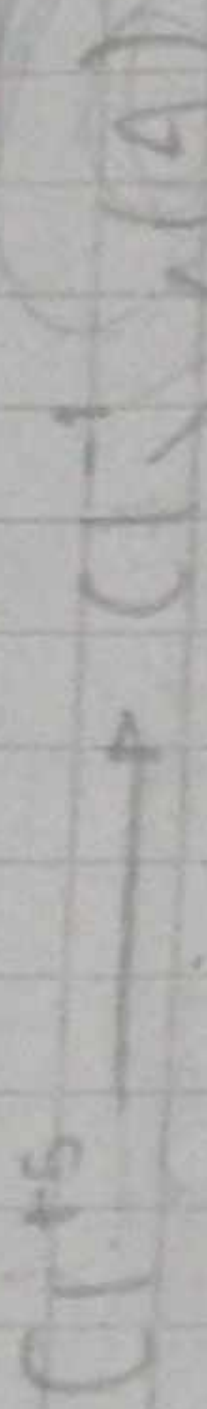


1 Sn 1

4 N 4

4 H 4

12 O 12

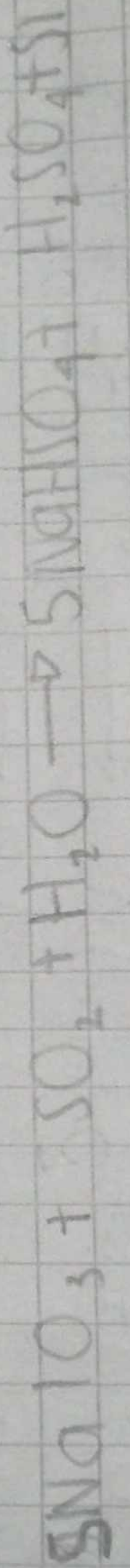
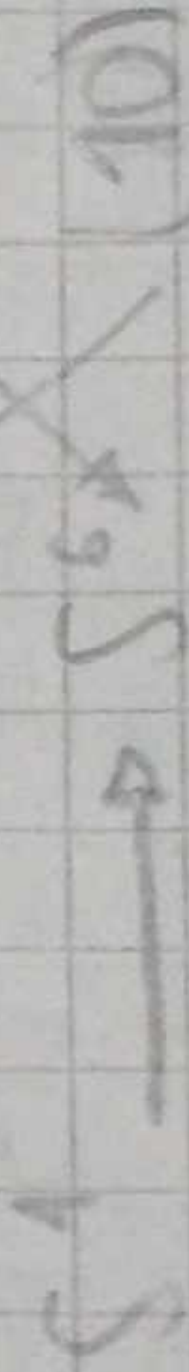
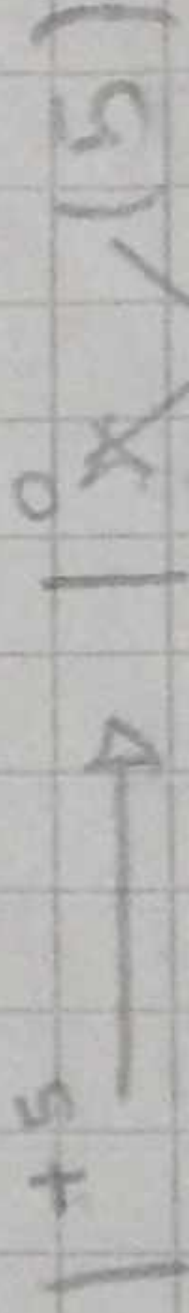
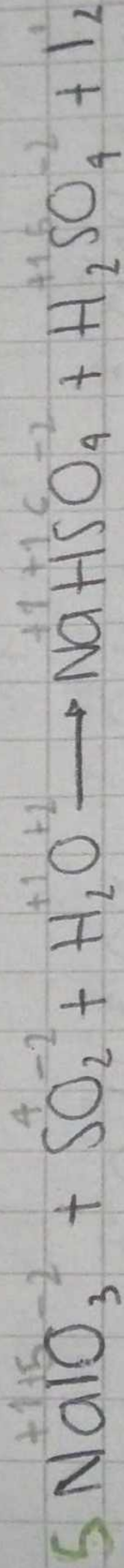


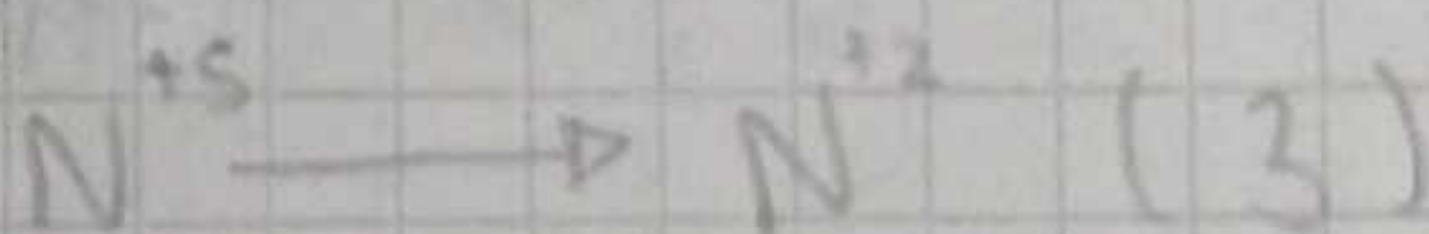
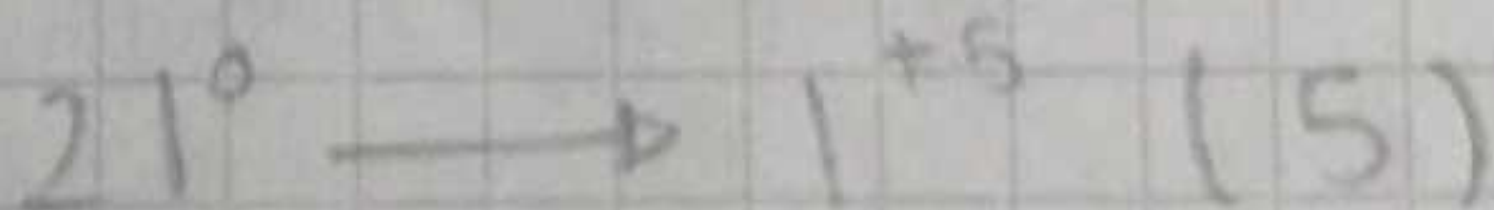
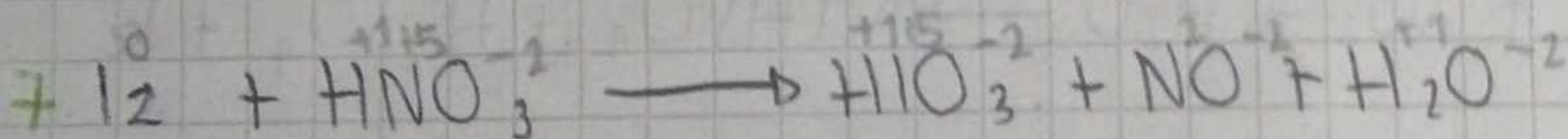
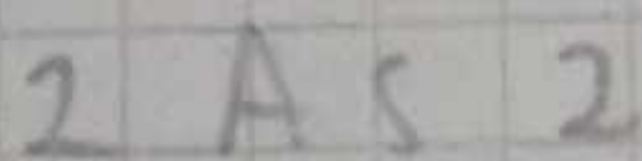
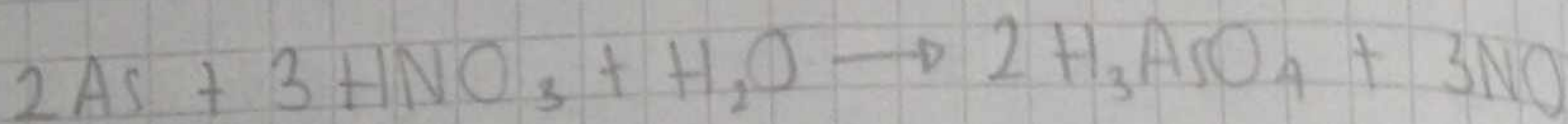
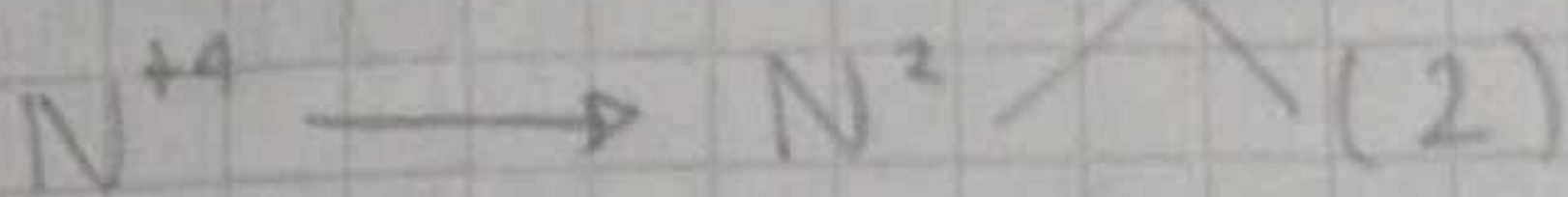
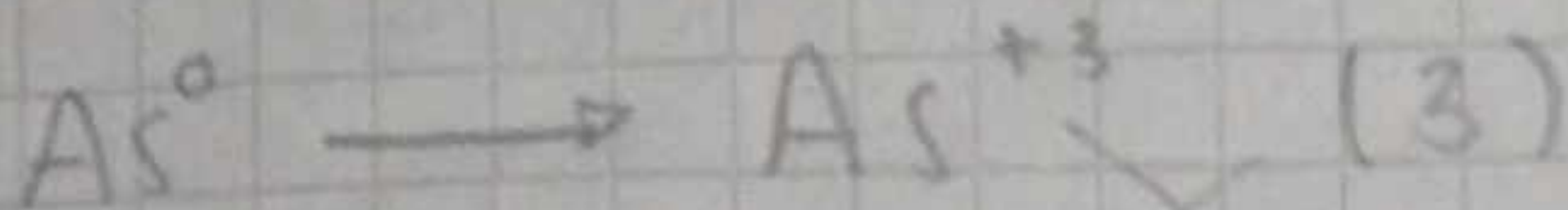
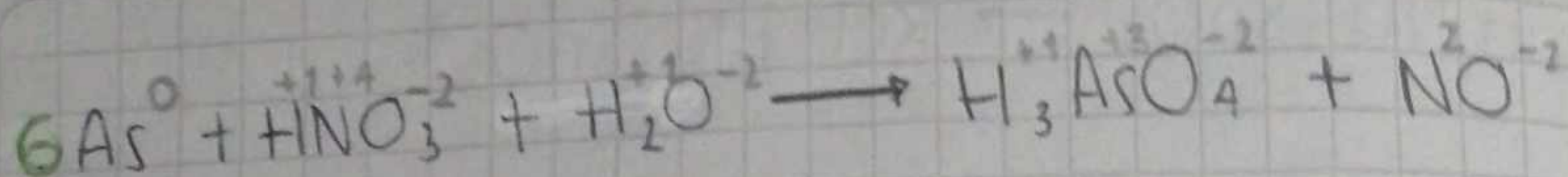
8 Na 8

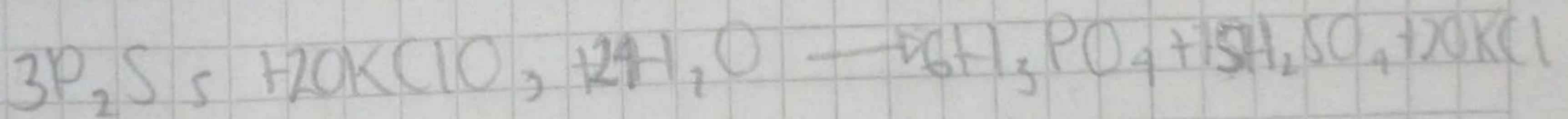
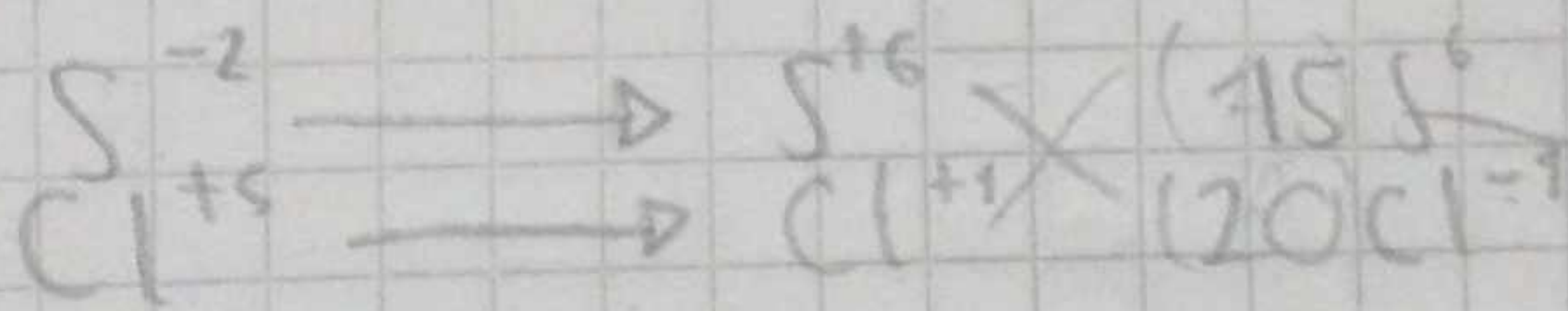
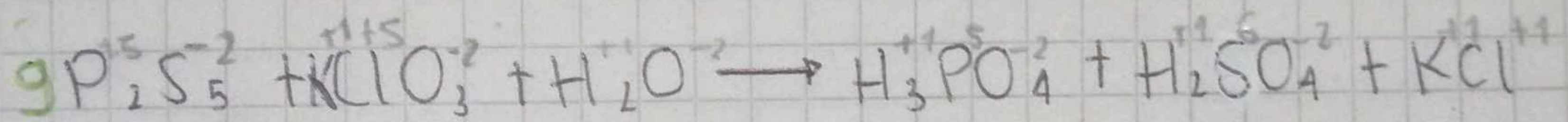
2 K 2

2 Cl 2

12 O 12







6 P 6

15 S 15

20 K 20

20 Cl 20

