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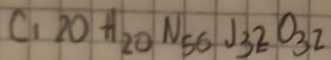
$$C = \frac{98 \text{ g}}{12 \text{ g/mol}} = 4 = \frac{48}{0.4} \quad 120$$

$$H = \frac{49}{1 \text{ g/mol}} = 4 = \frac{4}{0.4}$$

$$N = \frac{27.4 \text{ g}}{14 \text{ g/mol}} = 1.6 = \frac{21.4}{0.4} \quad 56$$

$$J = \frac{12.8 \text{ g}}{32 \text{ g/mol}} = 0.4 = \frac{12.8}{0.4} \quad 32$$

$$O = \frac{12.8 \text{ g}}{16 \text{ g/mol}} = 0.8 = \frac{2.8}{0.4}$$



$$C = 12 \times 120 = 1440$$

$$H = 1 \times 10 = 10$$

$$N = 14 \times 56 = 284$$

$$J = 32 \times 32 = 1024$$

$$O = 16 \times 32 = 512$$

