

1. Jolkeich

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$$C: 92,3$$

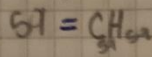
$$H: 7,7$$

$$\text{Mol: } 7,6910 = 69,13$$

$$\text{Mol: } 7,7 = 7,6916$$

$$\text{Mol C} = \frac{92,30}{12 \text{ g/mol}} = 7,6916$$

$$\text{Mol H} = \frac{7,7}{1 \text{ g/mol}} = 7,7$$



$$2 \text{ Na} = 32,4$$

$$\text{S} = 22,6$$

$$\text{O} = 45,1$$

$$- 1,4036$$

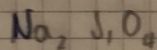
$$- 0,6963$$

$$- 2,813$$

$$\text{Mol Na} = \frac{32,40}{23 \text{ g/mol}} = 1,4036$$

$$\text{Mol S} = \frac{22,60}{32 \text{ g/mol}} = 0,69$$

$$\text{Mol O} = \frac{45,10}{16 \text{ g/mol}} = 2,813$$



$$3 \text{ C} = 98\%$$

$$\text{H} = 4\%$$

$$\text{N} = 22,4\%$$

$$\text{S} = 12,8\%$$

$$\text{O} = 12,8\%$$

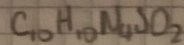
$$\text{C: } 98 \div 12 = 0,4 = 10$$

$$\text{H: } 4 \div 1 = 0,4 = 10$$

$$\text{N: } 22,4 \div 14 = 1,6 = 4$$

$$\text{S: } 12,8 \div 32 = 0,4 = 1$$

$$\text{O: } 12,8 \div 16 = 0,8 = 2$$



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