Solución Taller

① Un experimento arrojo los siguientes resultados:

C = 92.3 % = CH

H = 7.7 %

$C = \frac{92.3\,g}{12\,g/mol} = 7.691 - \frac{7.691}{7.7} = 1$

$H = \frac{7.7}{1\,g/mol} = 7.7 - \frac{7.7}{7.7} = 1$

② obten la formula empirica si hay=

Na: 32.4 % S: 22,5 % O: 45, 1 %

$Na = \frac{32.4\,g}{23\,g/mol} = 1.4 - \frac{1.4}{1.4} = 1$ $NaSO_2$

$S = \frac{22.5}{32.} = 1.4 - \frac{1.4}{1.4} = 1$

$O = \frac{45.1}{16} = 2.8 - \frac{2.8}{1.4} = 2$

③ Obten la formula empirica si hay:

C = 48% H = 4% N = 12.8% O = 12,8

$C = \frac{48gr}{12g} = 4 - \frac{4}{0.8} = 5$

$H = \frac{4}{1} = 4 - \frac{4}{0.8} = 5$ C_5H_5NO

$N = \frac{12.8}{14} = 1 - \frac{1}{0.8} = 1.25$

$O = \frac{12.8}{16} = 0.8 - \frac{08}{0.8} = 1$

④ Calcular la maja empirica $= N_2O$

N = 0.076 O = 0,181

$N = \frac{0.076}{14} = \frac{184}{} - \frac{184}{88} = 2$

$O = \frac{0,181}{16} = \frac{88}{} - \frac{88}{88} = 1$

(5) Determine la Formula empirica

Na = 21.6% Cl = 33.3%

Na = $\frac{21.6}{23} = 1 \quad \frac{1}{1} = 1$ NaCl

Cl = $\frac{33.3}{35} = 1 \quad \frac{1}{1} = 1$