

# MOLARIDAD

$$\text{Formula General} = \frac{\text{gramos}}{(\text{Pm}) (\text{Kg agua})}$$

$$\text{otra formula} = \frac{m}{(\text{Pm}) (v)}$$

## SOLUCIÓN EJERCICIOS

1. Alcohol etílico ( $\text{C}_2\text{H}_6\text{O}$ )

$$\begin{aligned} \text{masa} &= 82.5 \text{ g} \\ \text{volumen} &= 0.45 \\ \text{Pm} &= 46 \end{aligned}$$

$$\begin{aligned} \text{C} &= 24 \\ \text{H} &= 6 \\ \text{O} &= 16 \\ \hline &= 46 \end{aligned}$$

$$\begin{aligned} 82.5 \div 46 &= 1.793 \\ 1.793 \div 0.45 &= 3.985 \end{aligned}$$

$$\bullet \frac{82.5}{(46) \cdot (0.45)} = 20.7$$

$$\frac{82.5}{20.7} = 3.985$$

2) molaridad de la sustancia

$$\text{masa} = 4,78$$

$$V = 7$$

$$\frac{4,78}{7} = 0,682$$

3) Calcule la molaridad

$$M = 3,65$$

$$V = 2,00$$

$$PM = 36$$

$$H \times 1 = 1$$

$$Cl \times 1 = \frac{35}{36}$$

$$3,65 \div 36 = 0,10$$

$$0,10 \div 2,00 = 0,05$$

$$\bullet \frac{3,65}{(36)(2,00)} = \frac{3,65}{72} = 0,05$$

$$\begin{aligned} \uparrow \text{masa} &= 49,09 \\ \text{Volumen} &= 0,25 \\ PM &= 48 \end{aligned}$$

$$\begin{aligned} 2 \times 1 &= 2 \\ 1 \times 32 &= 32 \\ 4 \times 16 &= \frac{64}{78} \end{aligned}$$

$$49,09 \div 78 = 0,60$$

$$0,60 \div 0,25 = 2$$

$$\bullet \frac{49,09}{(78)(0,25)} = \frac{49,09}{24,5} = 2$$

