

... (OH-) es de

Solución

① $pOH = 14 - 9.6 = 4.4$ $(H^+) = 2.511 \times 10^{-10} M$

$(OH^-) = 3.981 \times 10^{-5} M$

② $pH = 14 - 2.50 = 11.5$ $(H^+) = 3.162 \times 10^{-12} M$

$(OH^-) = 0.003 M$

③ $pH = 5.6$

$(H^+) = 2.4 \times 10^{-6} M$

$pOH = 8.4$

$(OH^-) = 3.981 \times 10^{-9} M$

④ $pH = 8.7$

$(H^+) = 1.995 \times 10^{-9} M$

$pOH = 5.3$

$(OH^-) = 4.45 \times 10^{-6} M$