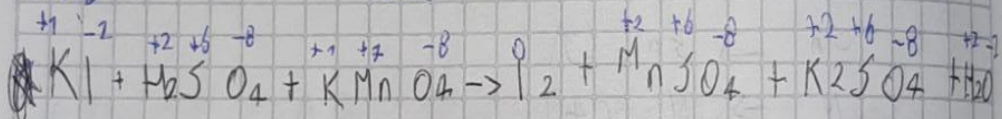


Tarea

- a) $KI + H_2SO_4 + KMnO_4 \rightarrow I_2 + MnSO_4 + K_2SO_4$
b) $KIO_3 + Al + HCl \rightarrow I_2 + AlCl_3 + KCl$
c) $PbO_2 + Pb + H_2SO_4 \rightarrow PbSO_4 + H_2O$
d) $H_2S + HNO_3 \rightarrow S + NO + H_2O$
e) $KCrO_2 + KClO + KOH \rightarrow K_2CrO_4 + KCl + H_2O$

12a)

solución

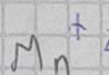


$$O = -2$$

$$H = +1$$

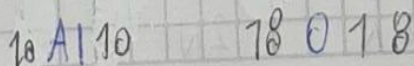
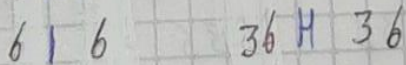
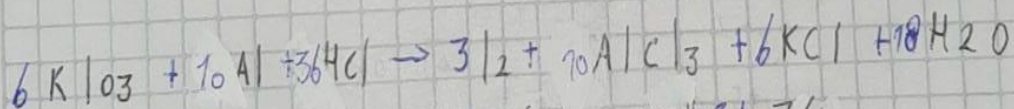
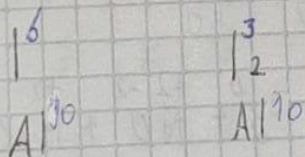
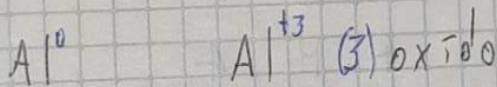
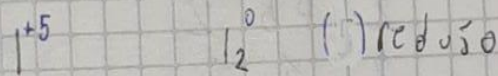
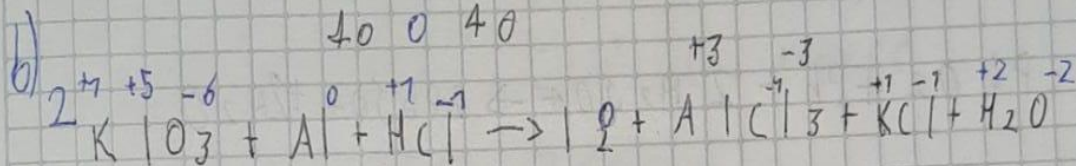
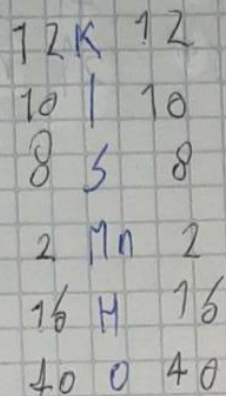
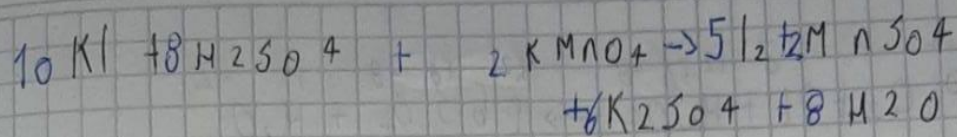


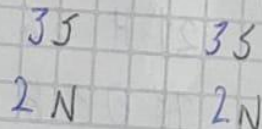
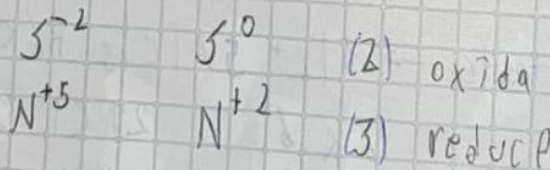
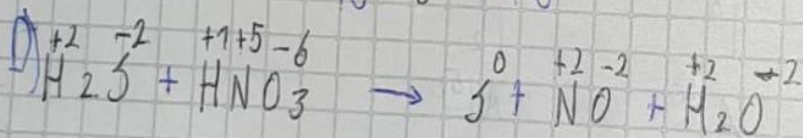
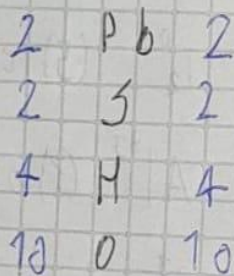
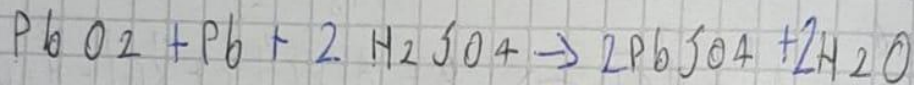
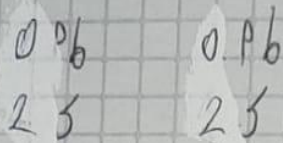
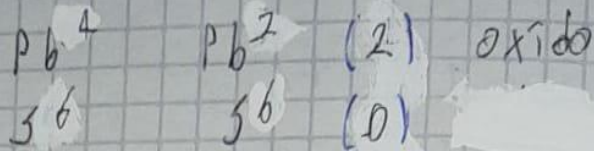
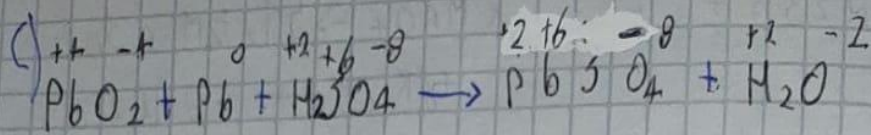
(2) oxidado

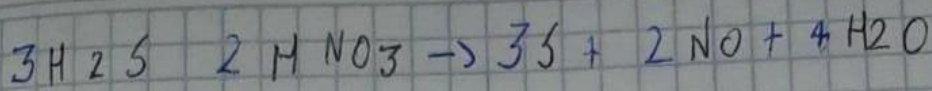


(5) reducido







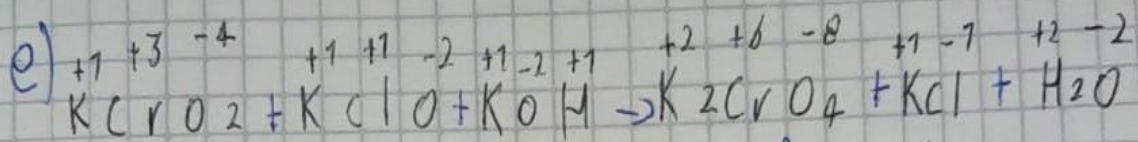


3 S 3

2 N 2

8 H 8

6 O 6

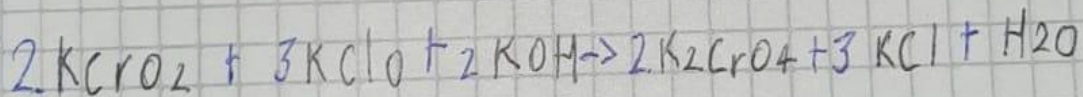


Cr⁺³ Cr⁺⁶ (3) oxido

Cl⁺¹ Cl⁻¹ (2) reduce

2 Cr 2 Cr

3 Cl 3 Cl



7 K 7

2 Cr 2

3 Cl 3

2 H 2

9 O 9