

$$1 \quad C = \frac{98g}{12g/mol} = 8.1667 = \frac{48}{0.4} = 120$$

$$H = \frac{9g}{1g/mol} = 9 = \frac{4}{0.4}$$

$$N = \frac{22.4g}{14g/mol} = 1.6 = \frac{21.4}{0.4} = 56$$

$$S = \frac{12.8g}{32g/mol} = 0.4 = \frac{12.8}{0.4} = 32$$

$$O = \frac{12.8g}{16g/mol} = 0.8 = \frac{12.8}{0.4}$$

$C_{120}H_{10}N_{56}S_{32}O_{32}$

$$C = 12 \times 120 = 1440$$

$$H = 1 \times 10 = 10$$

$$N = 14 \times 56 = 784$$

$$S = 32 \times 32 = 1024$$

$$O = 16 \times 32 = 512$$

$C_{120}H_{10}N_{56}S_{32}O_{32} = C_{35}H_3N_{16}S_4O_4$