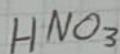


Ejercicios de molaridad

1. Calcular la molaridad de una disolución de 95 gramos de ácido nítrico (HNO_3) en 250 ml de agua



$$m = \frac{\text{gramos}}{(\text{Pm}) (\text{Kg})}$$

$$\text{Masa} = 95\text{g}$$

$$\text{volumen} = 0,25$$

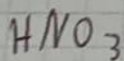
$$Pm = 63$$

$$H = 1 \times 1 = 1$$

$$N = 1 \times 14 = 14$$

$$O = 3 \times 16 = \frac{48}{63}$$

$$m = \frac{95}{(63) (0,25)} = \frac{95}{1,575} = 60,377\text{ m}$$



$$\text{Masa} = 95\text{g}$$

$$\text{volumen} = 0,25$$

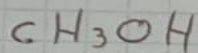
$$m = \frac{m}{\text{Kg}}$$

- (d) common, usual, customary
3. Based on information in the passage, it can be understood that is

$$\frac{95g}{63g/mol} = 1,507 \text{ molar}$$

$$\frac{1,507 m}{0,025} = 60,317 m$$

2. Calcular la molaridad de metanol (CH_3OH) en una disolución 15g donde el disolvente son 50 gramos de agua



$$m = \frac{\text{gramos}}{(P_m) (Kg)}$$

$$\text{Masa} = 15g$$

$$\text{volumen} = 0,05$$

$$P_m = 32$$

$$C = 1 \times 12 = 12$$

$$H = 3 \times 1 = 3$$

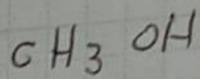
$$O = 1 \times 16 = 16$$

$$H = 1 \times 1 = 1$$

$$\underline{\quad}$$
$$32$$

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$$n = \frac{15}{(32)(0,05)} \quad \frac{15}{16} = 0,9375 \text{ m}$$



Masa = 15g

Volume = 0,05

$\rho_m = 32$

$$m = \frac{m}{kg}$$

$$\frac{15g}{32g/mol} = 0,468 \text{ moles}$$

$$\frac{0,468}{0,05} = 9,375 \text{ m}$$