

$$1) C = 40\%$$

$$H = 6,7\%$$

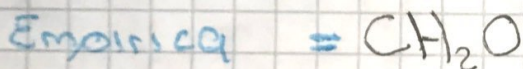
$$O = 53,3\%$$

$$\text{Masa} = 90\text{g}$$

$$C = \frac{40}{12} = 3,333 \quad \text{mol } 3,333 / 3,331 = 1$$

$$H = \frac{6,7}{1} = 6,7 \quad \text{mol } 6,7 / 3,331 = 2$$

$$O = \frac{53,3}{16} = 3,331 \quad \text{mol } 3,331 / 3,331 = 1$$



$$C = 12 \cdot 1 = 12$$

$$H = 1 \cdot 2 = 2$$

$$O = 16 \cdot 1 = 16$$

$$\underline{\quad 30}$$

$$n = 90\text{g} / 30 = 3$$



$$2) C = 37,8\%$$

$$H = 6,3\%$$

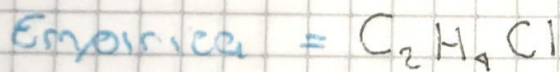
$$Cl = 55,8\%$$

$$\text{Masa} = 127g$$

$$C = \frac{37,8}{12} = 3,15 \quad \text{mol} \quad 3,15 / 1,571 = 2$$

$$H = \frac{6,3}{1} = 6,3 \quad \text{mol} \quad 6,3 / 1,571 = 4$$

$$Cl = \frac{55,8}{35,5} = 1,571 \quad \text{mol} \quad 1,571 / 1,571 = 1$$



$$C = 12 \cdot 2 = 24$$

$$H = 1 \cdot 4 = 4$$

$$Cl = 35,5 \cdot 1 = 35,5$$

$$\underline{63,5}$$

$$N = 127g / 63,5 = 2$$

