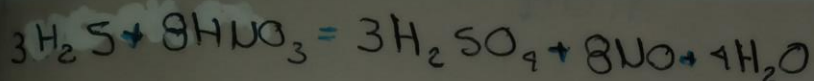
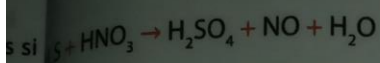
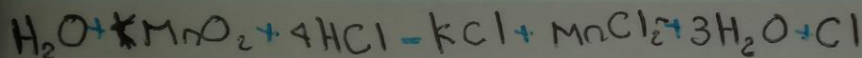
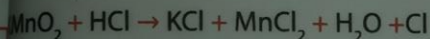


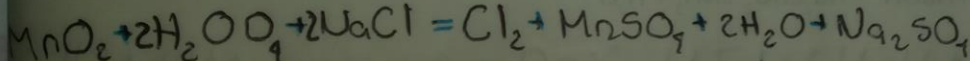
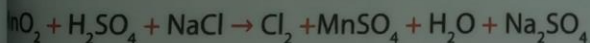
1. Balancea por óxido-reducción las siguientes ecuaciones químicas, teniendo en cuenta los números de oxidación y plantea semirreacciones para cada una, indica quien se oxida y quien se reduce.



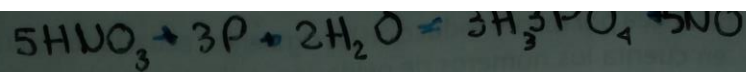
3	S	3
8	N	8
14	H	14
24	O	24



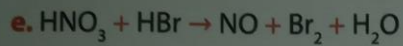
1	K	1
1	Mn	1
4	Cl	4
6	H	6
3	O	3



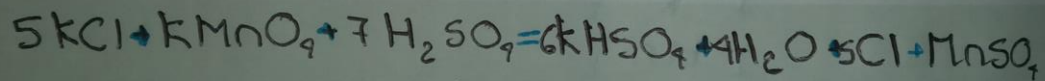
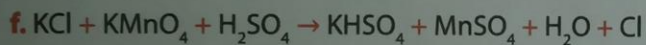
1	Mn	1
2	S	2
2	Na	2
2	Cl	2
4	H	4
10	O	10



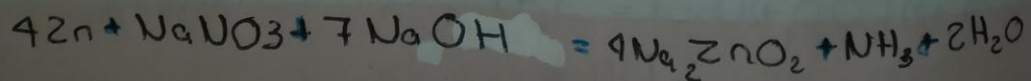
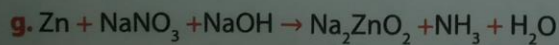
5 N 5
3 P 3
9 H 9
17 O 17



2 N 2
6 Br 6
8 H 8
6 O 6



6 K 6
5 Cl 5
1 Mn 1
14 H 14
32 O 32



4 Zn 4
8 Na 8
1 N 1
7 H 7
10 O 10

