

1) Soluto = 55g Solvente = 239g

$$\begin{array}{r}
 S \ 55 \text{ gr} \\
 O \ 239 \text{ gr} \\
 \hline
 294 \text{ gr}
 \end{array}
 \quad
 \% \frac{55}{294} \times 100 = 18,7\%$$

2) Soluto = 2,3kg Solvente = 6L

$$\begin{array}{r}
 S = 2,3 \text{ kg} \\
 O = 6 \text{ L} \\
 \hline
 8,3 \text{ L}
 \end{array}
 \quad
 \frac{2,3 \text{ kg}}{8,3 \text{ L}} \cdot 100 = 27\%$$

~~SO~~

1) Soluto: 13g Solución: 110g

$$\frac{13g}{110g} \cdot 100 = 11.818\%$$

2) Soluto = 10g Solución = 123

$$\frac{10g}{123g} \cdot 100 = 8.130\%$$

3) Soluto = 2,09ml Solvente = 15ml

$$S = 2,09ml$$

$$D = 15ml$$

$$\frac{2,09}{17,09}$$

$$\frac{2,09}{17,09} \cdot 100 = 11,9\%$$

4) Soluto 2,09 Solvente: 11ml

$$S = 2,09ml$$

$$D = 11ml$$

$$\frac{2,09}{13,09}$$

$$\frac{2,09}{13,09} \cdot 100 = 15,96\%$$