

$$\frac{0.181 \text{ g}}{16 \text{ g/mol}} = 0.0113$$

$$S) 21.6\%$$

$$Cl) 33.3\%$$

$$O) 45.1\%$$

$$\text{Mol } 0.675 \times 0.675 = 0.453$$

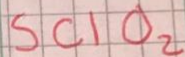
$$\text{Mol } 0.9314 \times 0.675 = 0.642$$

$$\text{Mol } 2.81875 \times 0.675 = 1.902$$

$$\text{Mol } S \frac{21.6 \text{ g}}{32 \text{ g/mol}} = 0.675$$

$$\text{Mol } Cl \frac{33.3 \text{ g}}{35.5 \text{ g/mol}} = 0.9314$$

$$\text{Mol } O \frac{45.1 \text{ g}}{16 \text{ g/mol}} = 2.81875$$



Correction

$$1) 7.6 \div 7.6 = 1 \quad CH$$

$$7.7 \div 7.6 = 1$$

$$1.6 \div 0.4 = 4$$

$$0.4 \div 0.4 = 1$$

$$2) 1.4 \div 0.6 = 2 \quad Na_2SO_4$$

$$0.6 \div 0.6 = 1$$

$$2.8 \div 0.6 = 4$$

$$12.8 \div 0.4 = 32$$

$$4) 0.5 \div 0.5 = 1 \quad NO_2$$

$$0.1 \div 0.05 = 2$$

$$3) 4 \div 0.4 = 10 \quad C_{10}H_{16}N_4SO_{32}$$

$$4 \div 0.4 = 10$$

$$5) 0.6 \div 0.6 = 1 \quad SClO_2$$

$$0.9 \div 0.6 = 1.5$$

$$2.8 \div 0.6 = 4$$