

Actividad

1) C = 92,3%

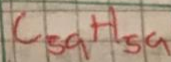
H = 7,7%

$$\text{Mol C} = \frac{92,3\text{g}}{12\text{g/mol}} = 7,6916$$

$$\text{Mol H} = \frac{7,7\text{g}}{1\text{g/mol}} = 7,7$$

$$\text{Mol } 7,6916 \times 7,6916 = 59,13$$

$$\text{Mol } 7,7 \times 7,6916 = 59,21$$



2) Na 32,4%

S 22,5%

O 45,1%

$$\text{Mol Na} = \frac{32,4\text{g}}{23\text{g/mol}} = 1,4086$$

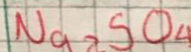
$$\text{Mol S} = \frac{22,5\text{g}}{32\text{g/mol}} = 0,696875$$

$$\text{Mol O} = \frac{45,1\text{g}}{16\text{g/mol}} = 2,81875$$

$$\text{Mol } 1,4086 \times 0,696875 = 2$$

$$\text{Mol } 0,696875 \times 0,696875 = 1$$

$$\text{Mol } 2,81875 \times 0,696875 = 4$$



3) C 48%

H 4%

N 22,4%

S 12,8%

O 12,8%

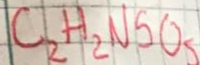
$$\text{Mol } 4 \times 0,4 = 1,6$$

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$$\text{Mol } 1,6 \times 0,4 = 0,64$$

$$\text{Mol } 0,4 \times 0,4 = 0,16$$

$$\text{Mol } 12,8 \times 0,4 = 5,12$$



$$\text{Mol C} = \frac{48\text{g}}{12\text{g/mol}} = 4 \quad \text{Mol S} = \frac{12,8\text{g}}{32\text{g/mol}} = 0,4$$

$$\text{Mol H} = \frac{4\text{g}}{1\text{g/mol}} = 4 \quad \text{Mol O} = \frac{12,8\text{g}}{16\text{g/mol}} = 0,8$$

$$\text{Mol N} = \frac{22,4\text{g}}{14\text{g/mol}} = 1,6$$

4) ~~_____~~

N $\frac{0,034g}{14g/Mol} = 0,0024$

O $\frac{0,181g}{16g/Mol} = 0,0113$

Mol $0,0036 \times 0,0056 = 0,0024$

Mol $0,0113 \times 0,0036 = 0,00063$

NO

5) S 21,6%

Cl 33,3%

O 45,1%

Mol S $\frac{21,6g}{32g/Mol} = 0,675$

Mol Cl $\frac{33,3g}{35g/Mol} = 0,9314$

Mol O $\frac{45,1g}{16g/Mol} = 2,81875$

Mol $0,675 \times 0,675 = 0,453$

Mol $0,9314 \times 0,675 = 0,642$

Mol $2,81875 \times 0,675 = 1,902$

SClO₂