

## Normalidad

1. Calcular la molaridad de una disolución de 95g de ácido nítrico ~~en~~ HNO<sub>3</sub> en 25 g de agua

$$\begin{aligned} m &= ? \\ \text{gramos} &= 95 \text{ g} \\ \text{agua} &= 25 / 1000 = 0.025 \\ \text{PH} &= 63 \text{ g/mol} \end{aligned}$$
$$\begin{aligned} \text{PH} &= 1 \times 1 \times 1 = 1 \\ \text{N} &= 1 \times 14 = 14 \\ \text{O} &= 3 \times 16 = 48 \\ &= 63 \text{ g/mol} \end{aligned}$$

$$m = \frac{95}{(63)(0.025)} = \frac{95}{1575} = 0.06$$

2. Calcular la molaridad de metanol (CH<sub>3</sub>OH) en una disolución de 15 g de disolvente son 50 gramos de agua

$$\begin{aligned} m &= ? \\ \text{gramos} &= 15 \text{ g} \\ \text{agua} &= 50 / 1000 = 0.05 \\ \text{PH} &= 32 \text{ g/mol} \end{aligned}$$
$$\begin{aligned} \text{PH} &= \text{C} = 1 \times 12 = 12 \\ \text{H} &= 4 \times 1 = 4 \\ \text{O} &= 1 \times 16 = 16 \\ &= 32 \text{ g/mol} \end{aligned}$$

$$m = \frac{15}{(32)(0.05)} = \frac{15}{16} = 0.9375$$